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NCERT Solutions for 12th Class Chemistry: Chapter 12-Aldehydes Ketones and Carboxylic Acids



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NCERT Solutions for 12th Class Chemistry: Chapter 12-Aldehydes Ketones and Carboxylic Acids

**Class 12: Chemistry Chapter 12 solutions. Complete Class 12
Chemistry Chapter 12 Notes.**

**NCERT Solutions for 12th Class Chemistry: Chapter
12-Aldehydes Ketones and Carboxylic Acids**

**NCERT 12th Chemistry Chapter 12, class 12 Chemistry chapter 12
solutions**

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EXERCISES (NCERT 12th Chemistry Chapter 12)

12.1. What is meant by the following terms? Give an example of the reaction in each case.

(i) Cyanohydrin

(ii) Acetal

(iii) Semicarbazone

(iv) Aldol

(v) Hemiacetal

(vi) Oxime

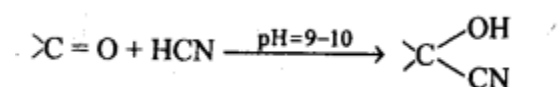
(vii) Ketal

(viii) Imine

(ix) 2,4-DNP derivative

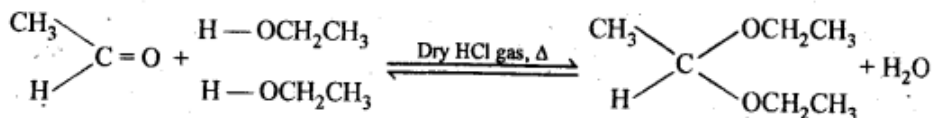
(x) Schiff's base.

Ans: (i) Cyanohydrin: gem-Hydroxynitriles, i.e., compounds possessing hydroxyl and cyano groups on the same carbon atom are called cyanohydrins. These are produced by addition of HCN to aldehydes or ketones in a weakly basic medium.

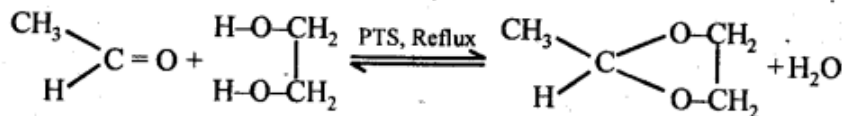


(ii) gem – Dialkoxy compounds in which the two alkoxy groups are present on the terminal carbon atom are called acetals. These are produced by the action of an aldehyde with two equivalents of a monohydric alcohol in presence of dry HCl gas.

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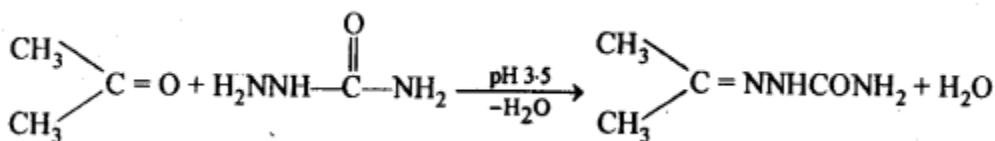


When dihydric alcohol is used cyclic acetal is formed

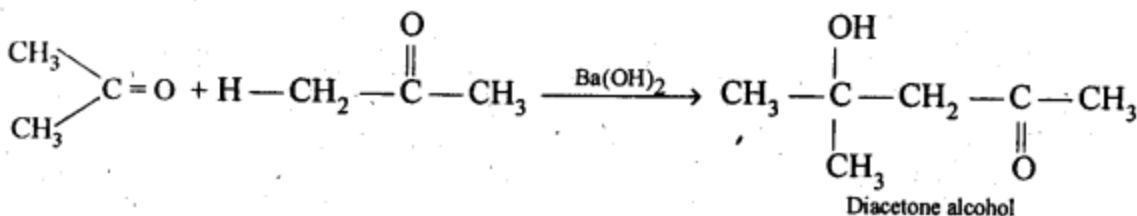
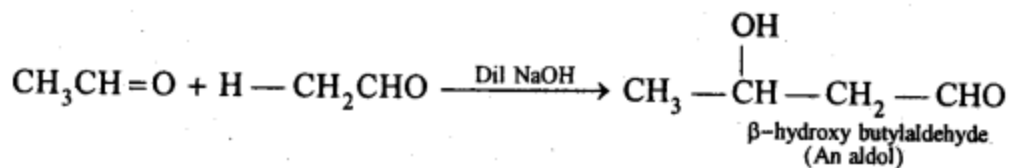


These are easily hydrolysed by dilute mineral acids to regenerate the original aldehydes. Therefore, these are used for the protection of aldehyde group in organic synthesis.

(iii) Semicarbazones are derivatives of aldehydes and ketones and are produced by action of semicarbazide on them in acidic medium.

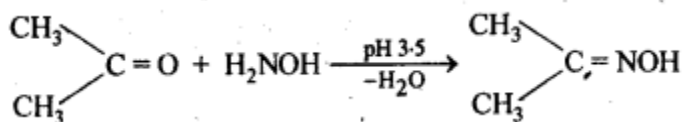


(iv) Aldols are β -hydroxy aldehydes or ketones and are produced by the condensation of two molecules of the same or one molecule each of two different aldehydes or ketones in presence of a dilute aqueous base. For example,

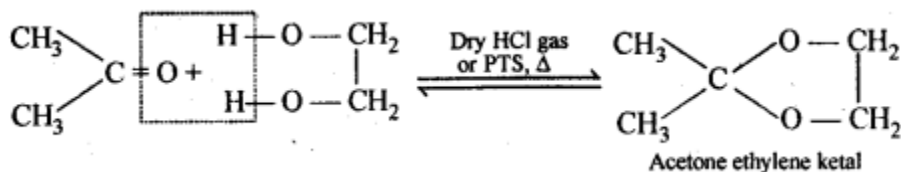


(v) gem – Alkoxyalcohols are called hemiacetals. These are produced by addition of one molecule of a monohydric alcohol to an aldehyde in presence of dry HCl gas.

(vi) Oximes are produced when aldehydes or ketones react with hydroxyl amine in weakly acidic medium.

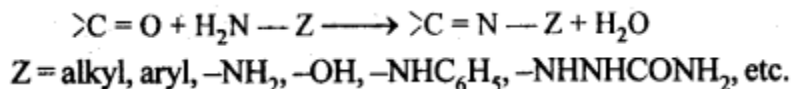


(vii) Ketals are produced when a ketone is heated with dihydric alcohols like ethylene glycol in presence of dry HCl gas or /3-toluene sulphonic acid (PTS).

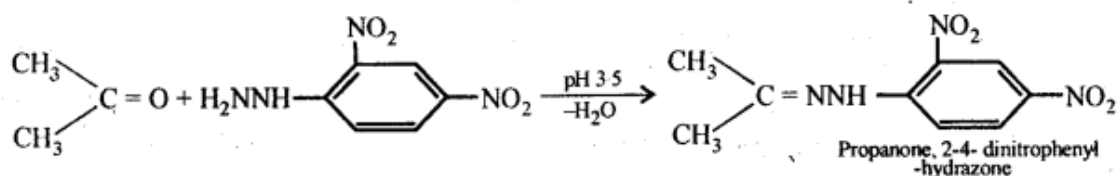


These are easily hydrolysed by dilute mineral acids to regenerate the original ketones. Therefore, ketals are used for protecting keto groups in organic synthesis.

(viii) Compounds containing $-C = N -$ group are called imines. These are produced when aldehydes and ketones react with ammonia derivatives.

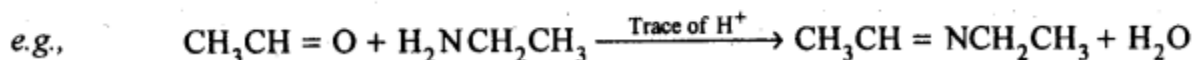
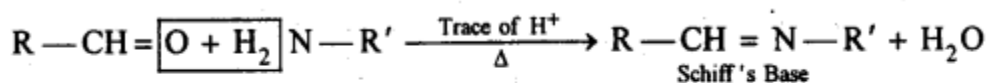


(ix) 2, 4-Dinitrophenyl hydrazone (i.e., 2,4-DNP derivatives) are produced when aldehydes or ketones react with 2,4-dinitrophenyl hydrazine in weakly acidic medium.



2, 4-DNP derivatives are used for identification and characterisation of aldehydes and ketones.

(x) Aldehydes and ketones react with primary aliphatic or aromatic amines to form azomethines or Schiff's bases.



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12.2. Name the following compounds according to IUPAC system of nomenclature:

(i) $CH_3CH(CH_3)-CH_2CH_2-CHO$

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Ans: (i) 4-Methyl pentanal

(ii) 6-Chloro-4-ethylhexan-3-one

(iii) But-2-en-1-al

(iv) Pentane-2,4-dione

(v) 3,3,5-Trimethyl-hexan-2-one

(vi) 3,3-Dimethyl butanoic acid

(vii) Benzene-1,4-dicarbaldehyde

12.3. Draw the structures of the following compounds.

(i) 3-Methylbutanal

(ii) p-Methylbenzaldehyde

(iii) 4-Chloropentan-2-one

(iv) p, p'-Dihydroxybenzophenone

(v) p-Nitropropiophenone

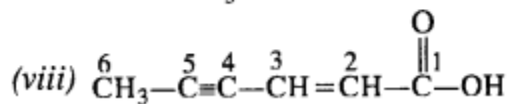
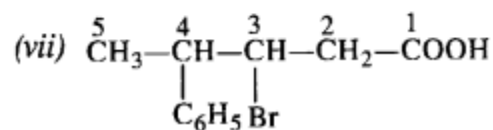
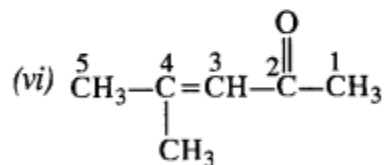
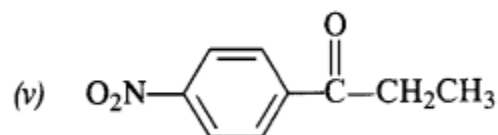
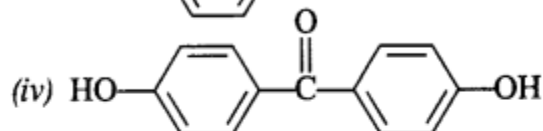
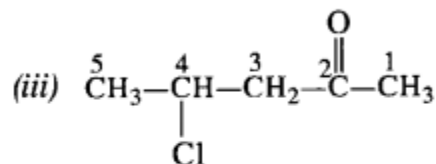
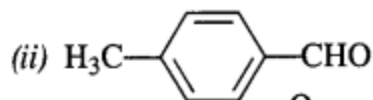
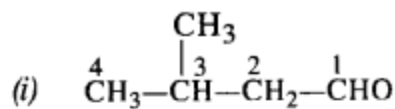
(vi) 4-Methylpent-3-en-2-one.

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(vii) 3-Bromo-4-phenylpentanoic acid

(viii) Hex-2-en-4-ynoic acid

Ans:



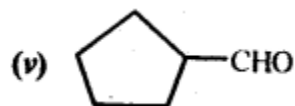
12.4. Write the IUPAC names of the following ketones and aldehydes. Wherever possible, give also common names.

(i) $\text{CH}_3\text{CO}(\text{CH}_2)_4\text{CH}_3$

(ii) $\text{CH}_3\text{CH}_2\text{CH BrCH}_2\text{CH}(\text{CH}_3)\text{CHO}$

(iii) $\text{CH}_3(\text{CH}_2)_5\text{CHO}$

(iv) $\text{Ph}-\text{CH}=\text{CH}-\text{CHO}$



(vi) Ph CO Ph

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Ans:

IUPAC name	Common name
(i) Heptan-2-one	—
(ii) 4-Bromo-2-methylhexanal	γ -Bromo- α -methyl caproaldehyde
(iii) Heptanal	—
(iv) 3-Phenylpropan-2-en-1-al	β -Phenyl acrolein
(v) cyclopentane carbaldehyde	—
(vi) Diphenylmethanone	Benzophenone

12.5. Draw structures of the following derivatives:

(i) The 2,4-dinitrophenylhydrazone of benzaldehyde

(ii) Cyclopropanone oxime

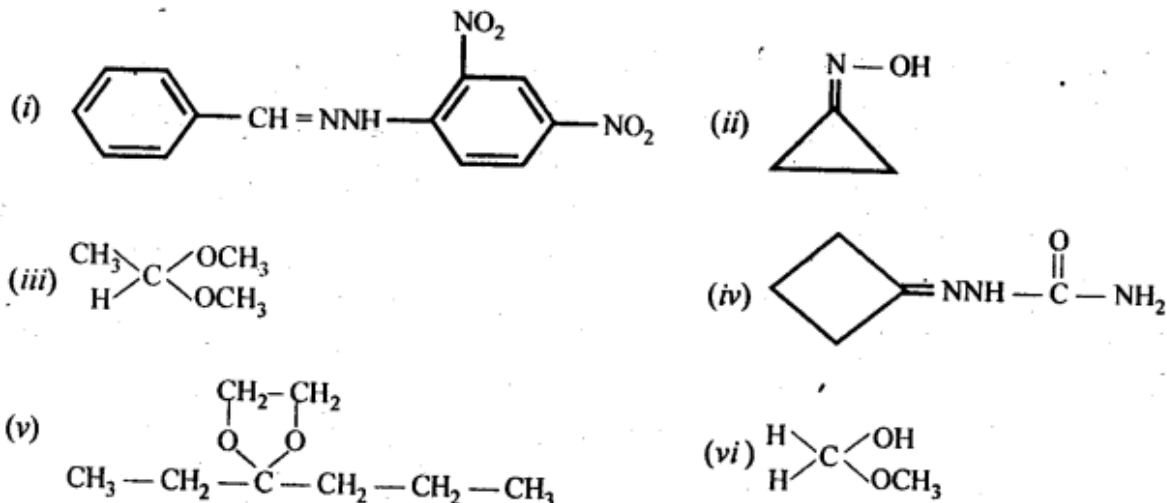
(iii) Acetaldehydedimethylacetal

(iv) The semicarbazone of cyclobutanone

(v) The ethylene ketal of hexan-3-one

(vi) The methyl hemiacetal of formaldehyde

Ans:



12.6. Predict the product when cyclohexanecarbaldehyde reacts with following reagents :

(i) $\text{C}_6\text{H}_5\text{MgBr}$ followed by H_3O^+

(ii) Tollen's reagent

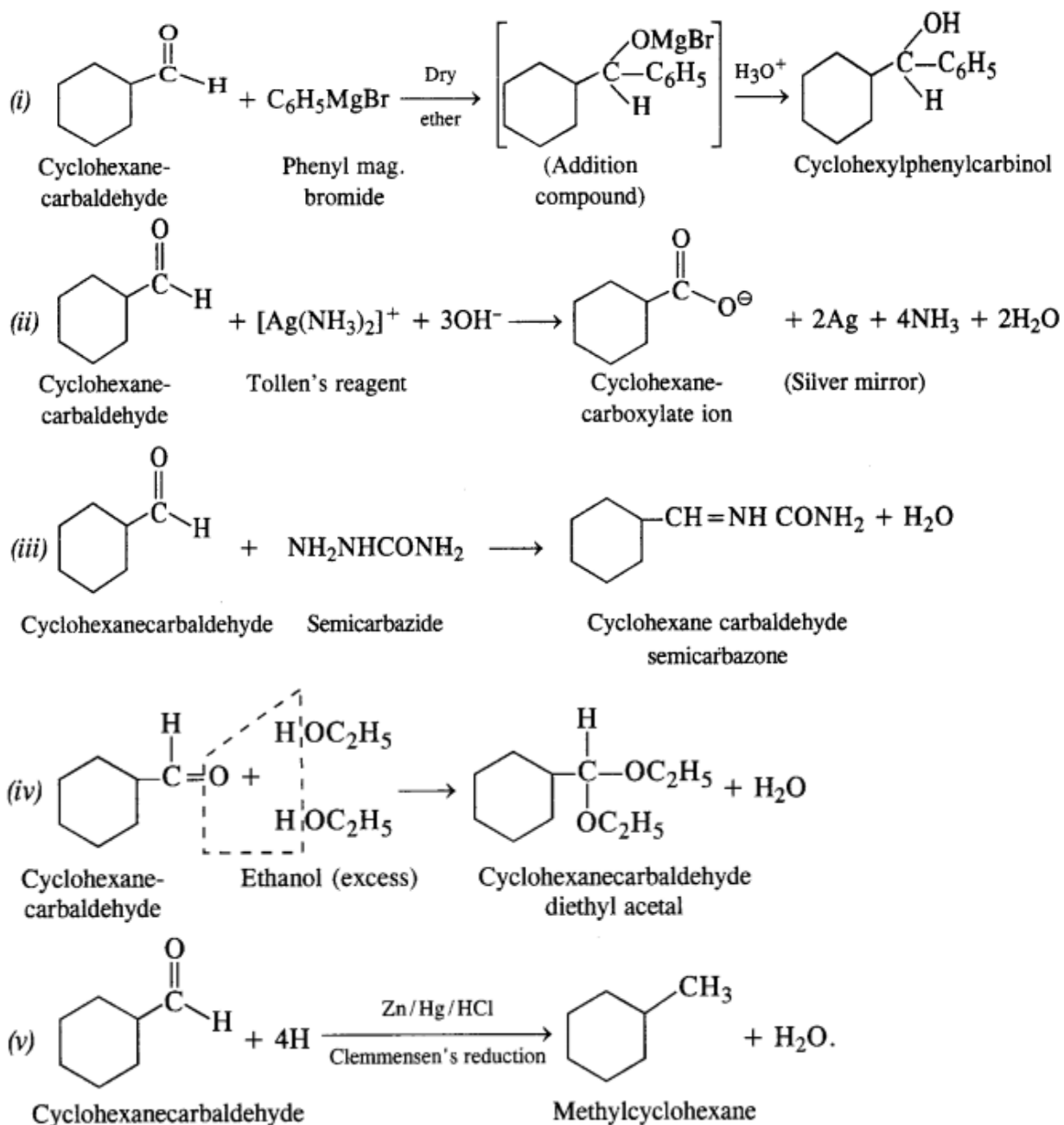
(iii) Semicarbazide in the weakly acidic medium

(iv) Excess of ethanol in the presence of acid

(v) Zinc amalgam and Cyclohexanecarbaldehyde Semicarbazide

Ans:

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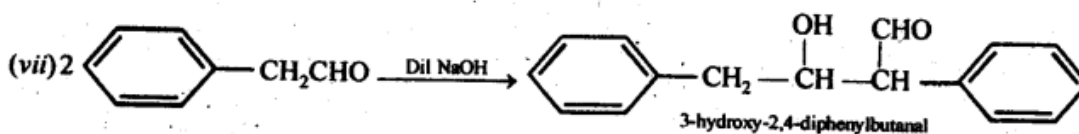
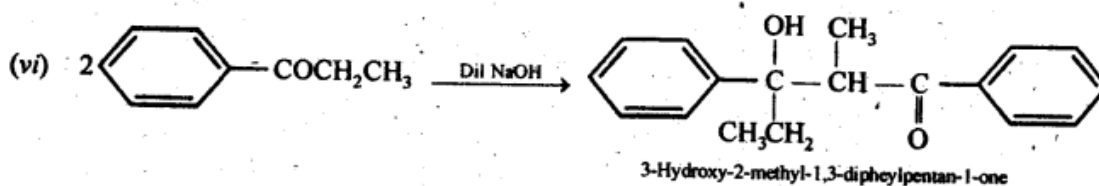
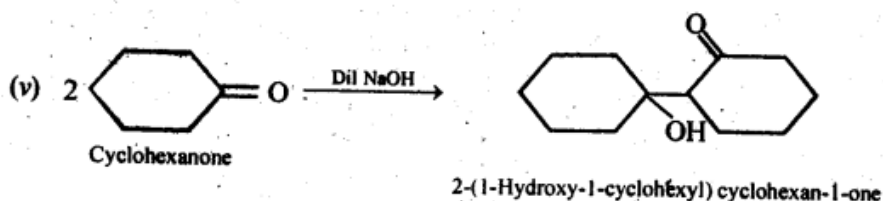
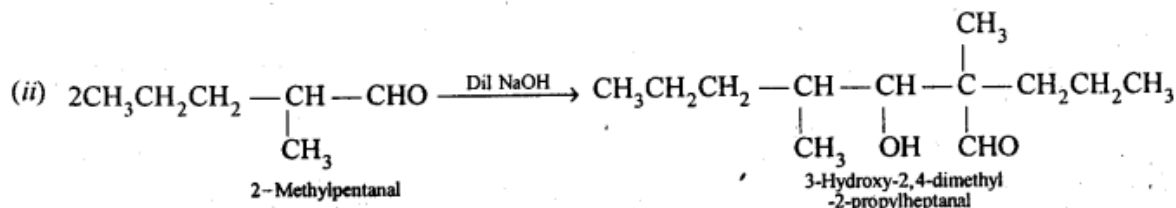
12.7. Which of the following compounds would undergo aldol condensation, which the Cannizzaro reaction and which neither? Write the structures of the expected products of aldol condensation and Cannizzaro reaction.

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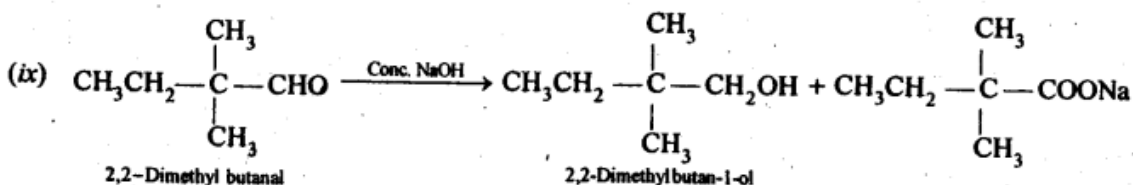
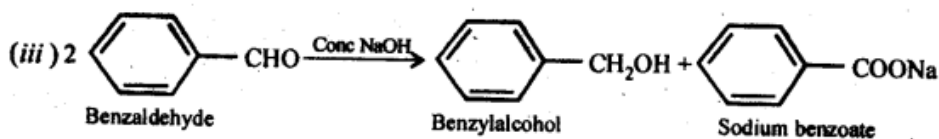
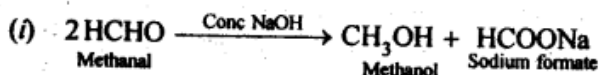
- (i) Methanal
- (ii) 2-Methylpentanal
- (iii) Benzaldehyde.
- (iv) Benzophenone
- (v) Cyclohexanone
- (vi) 1-Phenylpropanone
- (vii) Phenylacetaldehyde
- (viii) Butan-1-ol
- (ix) 2,2-Dimethylbutanal

Ans: 2-Methylpentanal, cyclohexanone, 1-phenylpropanone and phenylacetaldehyde contain one or more α -hydrogen and hence undergo aldol condensation. The reactions and the structures of the expected products are given below:

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Methanal, benzaldehyde and 2,2-dimethylbutanal do not contain α -hydrogen and hence undergo Cannizzaro reaction. The reactions and the structures of the expected products are given below:



Benzophenone (iv) is a ketone having no α -hydrogen while butan-1-ol (viii) is an alcohol. Both of these neither undergo aldol condensation nor cannizzaro reaction.

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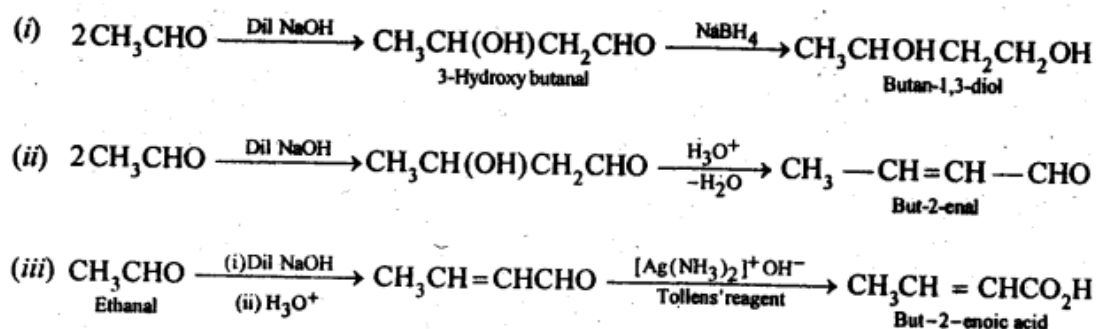
12.8. How will you convert ethanal into the following compounds?

(i) Butane-1,3-diol

(ii) But-2-enal

(iii) But-2-enoic acid

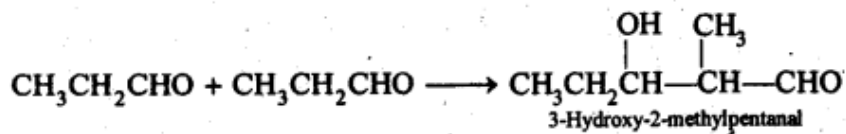
Ans:



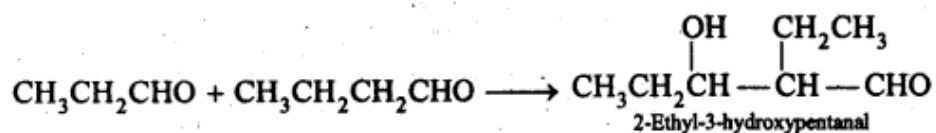
12.9. Write structural formulas and names of four possible aldol condensation products from propanal and butanal. In each case, indicate which aldehyde acts as nucleophile and which as electrophile.

Ans:

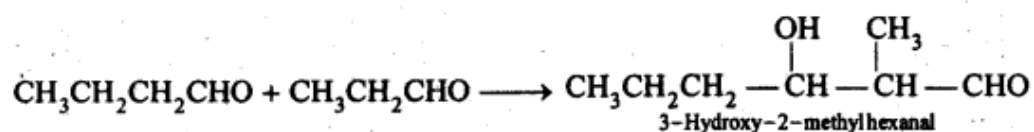
(a) Propanal acts as both nucleophile as well as electrophile.



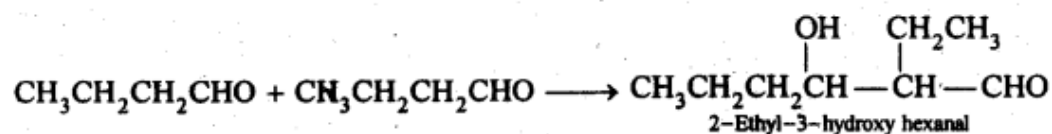
(b) Propanal as electrophile and butanal as nucleophile.



(c) Butanal as electrophile and propanal as nucleophile.



(d) Butanal acts as both nucleophile as well as an electrophile.



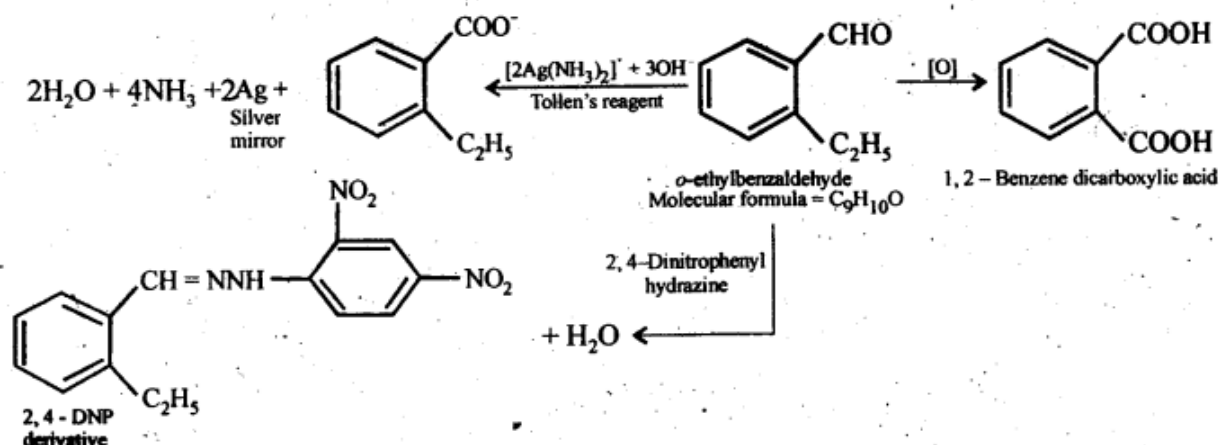
12.10. An organic compound with the molecular formula $\text{C}_9\text{H}_{10}\text{O}$ forms 2,4-DNP derivative, reduces Tollen's reagent, and undergoes Cannizzaro reaction. On vigorous oxidation, it gives 1,2-benzenedicarboxylic acid. Identify the compound.

Ans: Since the given compound with molecular formula $\text{C}_9\text{H}_{10}\text{O}$ forms a 2,4-DNP derivative and reduces Tollen's reagent, it must be an aldehyde. Since it undergoes Cannizzaro reaction, therefore, CHO group is directly attached to the benzene ring.

Since on vigorous oxidation, it gives 1, 2-benzene dicarboxylic acid, therefore, it must be an ortho- substituted benzaldehyde. The only o-substituted aromatic aldehyde having molecular

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formula $C_9H_{10}O$ is o-ethyl benzaldehyde. All the reactions can now be explained on the basis of this structure.



12.11. An organic compound (A) (molecular formula $C_8H_{16}O_2$) was hydrolysed with dilute sulphuric acid to give a carboxylic acid (B) and an alcohol (C). Oxidation of (C) with chromic acid produced (B). (C) on dehydration gives but-1-ene. Write equations for the reactions involved.

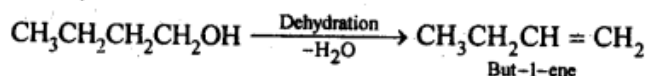
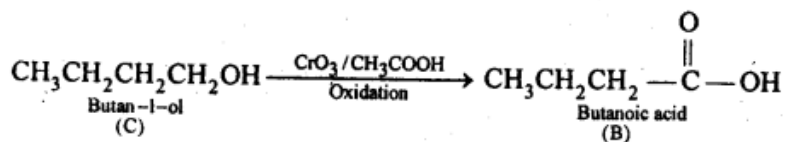
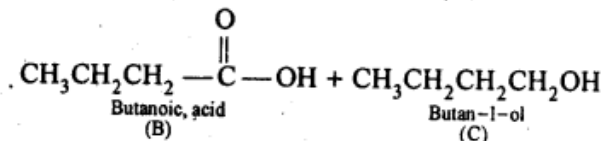
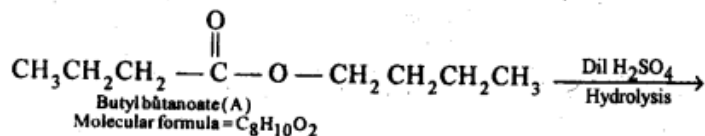
Ans: Since an ester A with molecular formula $C_8H_{16}O_2$ upon hydrolysis gives carboxylic acid B and the alcohol C and oxidation of C with chromic acid produces the acid B, therefore, both the carboxylic acid B and alcohol C must contain the same number of carbon atoms.

Further, since ester A contains eight carbon atoms, therefore, both the carboxylic acid B and the alcohol C must contain four carbon atoms each.

Since the alcohol C on dehydration gives but-1-ene, therefore, C must be a straight chain alcohol, i.e., butan-1-ol.

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If C is butan-1-ol, then the acid B must be butanoic acid and the ester A must be butyl butanoate. The chemical equations are as follows:



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12.12. Arrange the following in increasing order of the property indicated :

(i) Acetaldehyde, Acetone, Di tert. butyl ketone, Methyl tert. butyl ketone (reactivity towards HCN). (C.B.S.E. Sample Paper 2011, 2015, C.B.S.E. Delhi 2012)

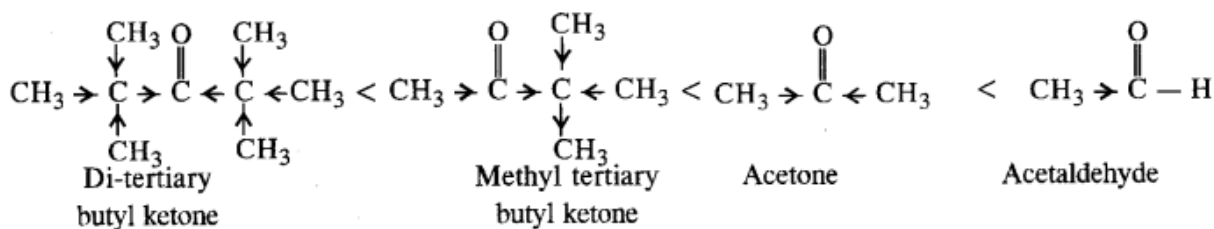
(ii) $\text{CH}_3\text{CH}_2\text{CH}(\text{Br})\text{COOH}$, $\text{CH}_3\text{CH}(\text{Br})\text{CH}_2\text{COOH}$, $(\text{CH}_3)_2\text{CHCOOH}$, $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}$ (acid strength) (C.B.S.E. Delhi 2008)

(iii) Benzoic acid, 4-Nitrobenzoic acid, 3, 5-Dinitrobenzoic acid, 4-Methoxybenzoic acid (acid strength) (C.B.S.E. Sample Paper 2011, 2015; C.B.S.E. Delhi 2012, C.B.S.E. Outside Delhi 2015, Rajasthan Board 2015)

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Ans: (i) Cyanohydrin derivatives are formed as a result of the reaction in which the nucleophile (CN^- ion) attacks the carbon atom of the carbonyl group. The order of reactivity

- decreases with increase in +I effect of the alkyl group.
- decreases with increase in steric hindrance due to the size as well as number of the alkyl groups. In the light of the above information, the decreasing order of reactivity is :

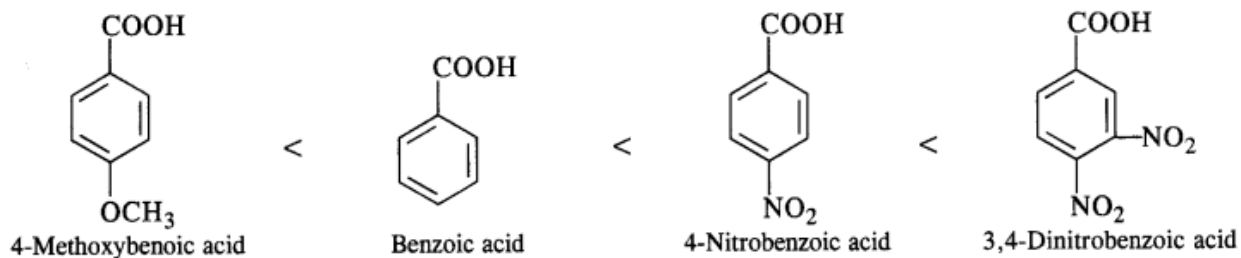


(ii) We know that alkyl group with +I effect decreases the acidic strength. The +I effect of isopropyl group is more than that of n-propyl group. Similarly, bromine (Br) with -I-effect increases the acidic strength. Closer its position in the carbon atom chain w.r.t., carboxyl (COOH) group, more will be its -I-effect and stronger will be the acid. In the light of this, the increasing order of acidic strength is :



(iii) We have learnt that the electron donating group (OCH_3) decreases the acidic strength of the benzoic acid. At the same time, the electron withdrawing group (NO_2) increases the same. Keeping this in mind, the increasing order of acidic strength is:

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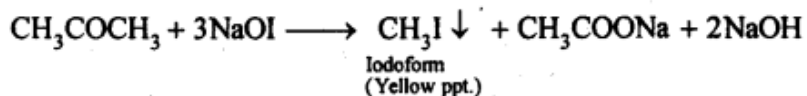
12.13. Give simple chemical tests to distinguish between the following pairs of compounds.

- (i) Propanal and Propanone**
- (ii) Acetophenone and Benzophenone**
- (iii) Phenol and Benzoic acid**
- (iv) Benzoic acid and Ethyl benzoate**
- (v) Pentan-2-one and Pentan-3-one**
- (vi) Benzaldehyde and Acetophenone.**
- (vii) Ethanal and Propanal**

Ans:

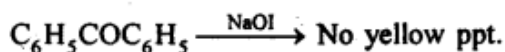
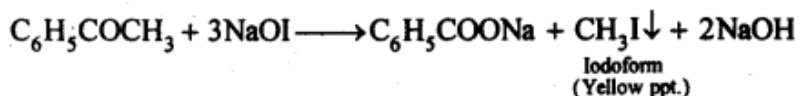
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- (i) Propanal and Propanone can be distinguished by iodoform test.



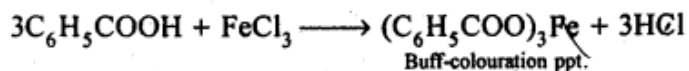
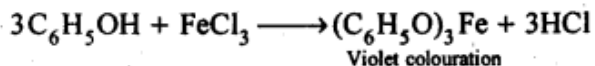
This test is given by aldehydes containing —COCH_3 group. Propanal does not have —COCH_3 group thus it does not give iodoform test

- (ii) Acetophenone and Benzophenone can be distinguished by iodoform test.



This test is given by aldehydes and ketones containing —COCH_3 group

- (iii) Phenol and Benzoic acid can be distinguished by FeCl_3 test

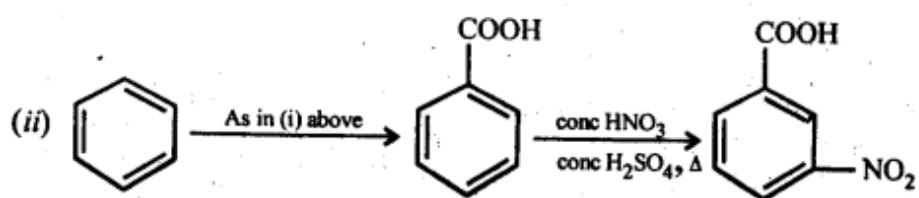
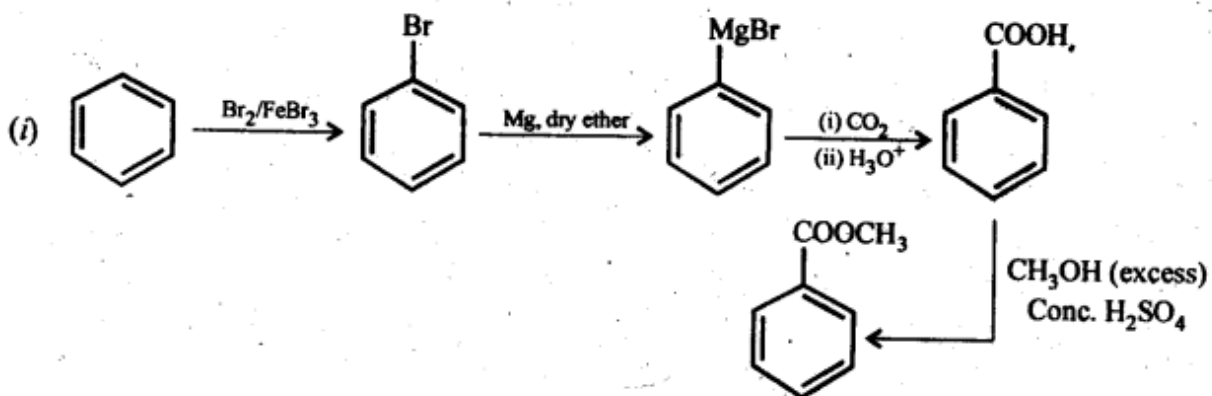


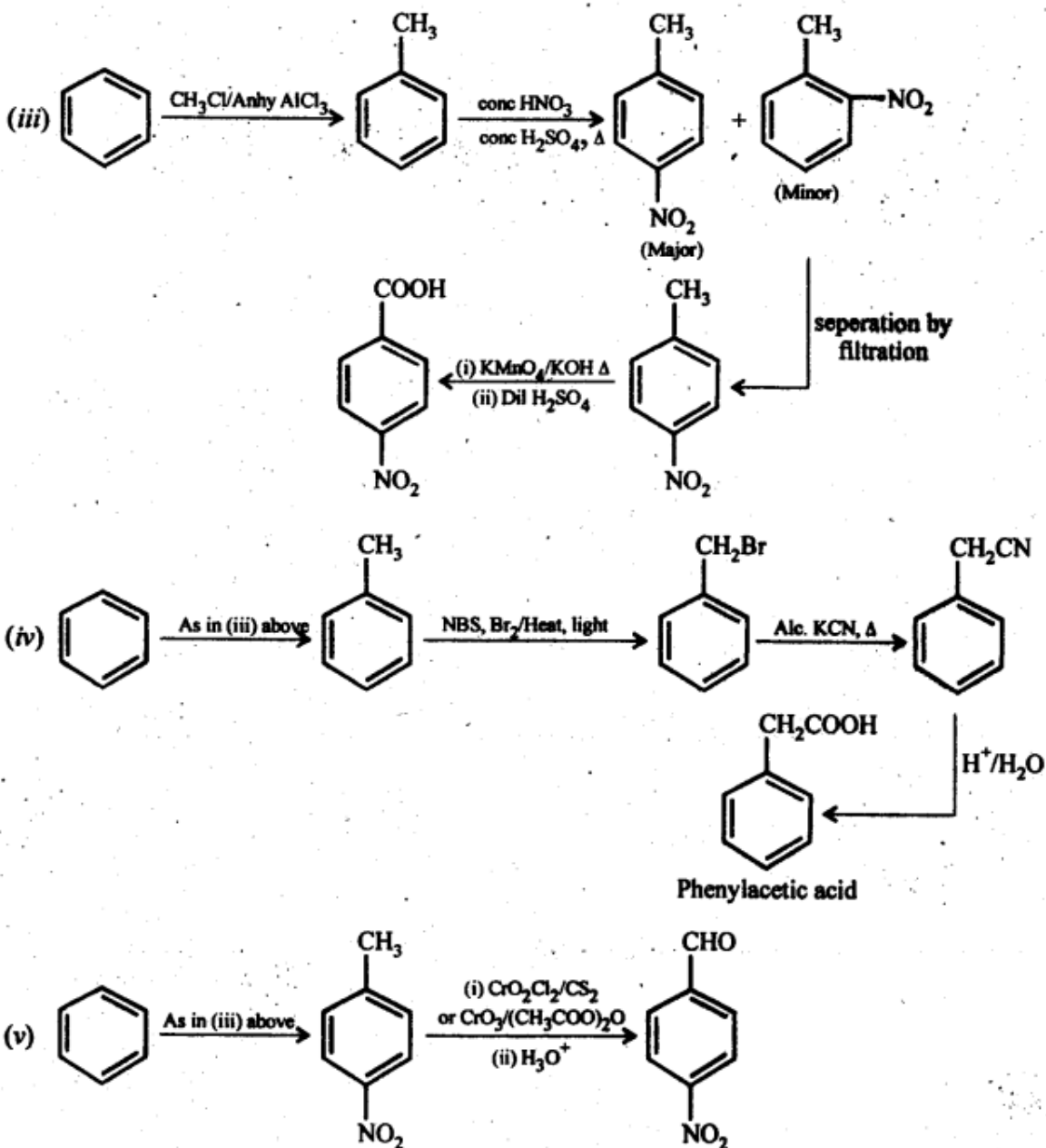
- (iv) Benzoic acid and Ethyl benzoate— By NaHCO_3 test

(v) p-nitrobenzaldehyde

Ans:

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12.15. How will you bring about the following conversions in not more than two steps?

(i) Propanone to Propene

(ii) Benzoic acid to Benzaldehyde

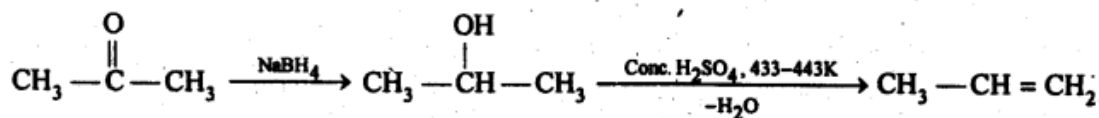
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- (iii) Ethanol to 3-Hydroxybutanal
- (iv) Benzene to m-Nitroacetophenone
- (v) Benzaldehyde to Benzophenone –
- (vi) Bromobenzene to 1-Phenylethanol
- (vii) Benzaldehyde to 3-Phenylpropan-1-ol.
- (viii) Benzaldehyde to α Hydroxyphenylacetic acid
- (ix) Benzoic acid to m-Nitrobenzyl alcohol

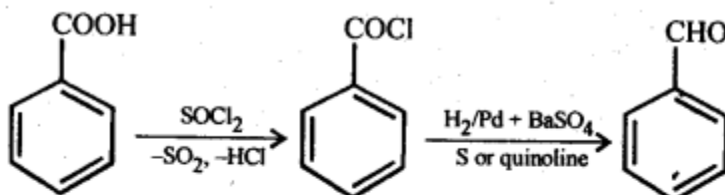
Ans:

<https://www.indcareer.com/schools/ncert-solutions-for-12th-class-chemistry-chapter-12-aldehydes-ketones-and-carboxylic-acids/>

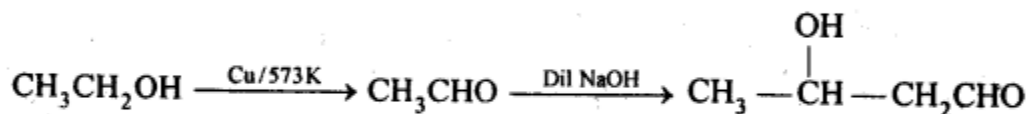
(i) Propanone to propene:



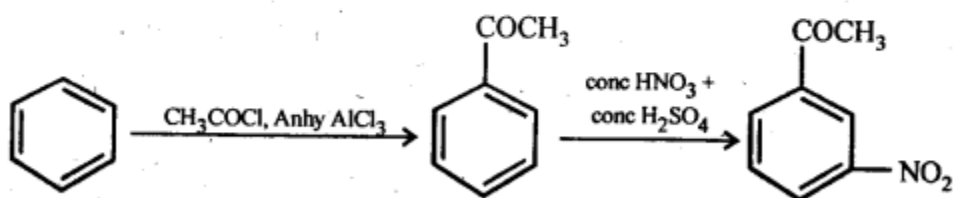
(ii) Benzoic acid to benzaldehyde:



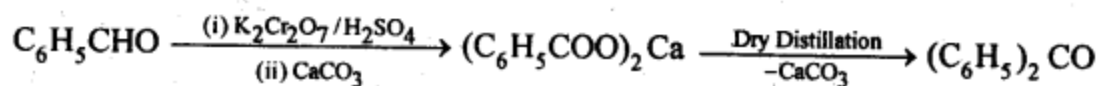
(iii) Ethanol to 3-hydroxy butanal:



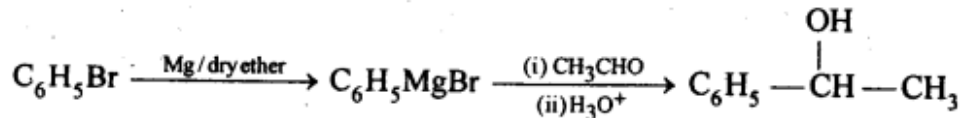
(iv) Benzene to *m*-nitroacetophenone:



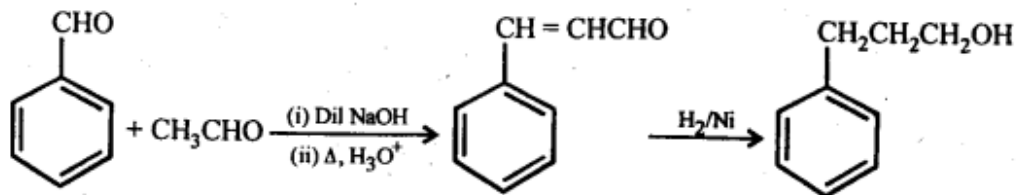
(v) Benzaldehyde to benzophenone:



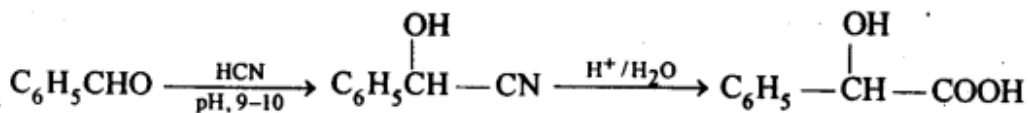
(vi) Bromobenzene to 1-phenylethanol



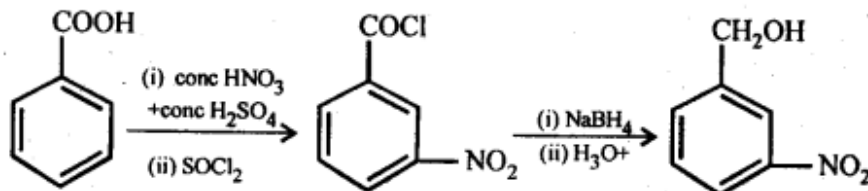
(vii) Benzaldehyde to 3-phenylpropan-1-ol



(viii) Benzaldehyde to α -hydroxyphenylacetic acid:



(ix) Benzoic acid to *m*-nitrobenzyl alcohol:



12.16. Describe the following:

(i) Acetylation

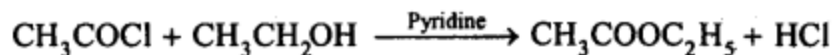
(ii) Cannizzaro reaction

(iii) Cross aldol condensation

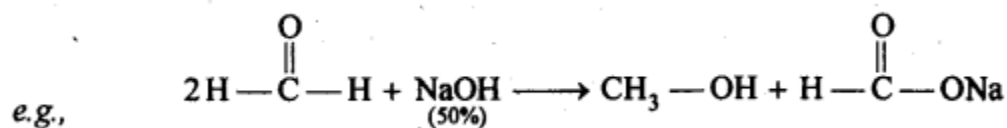
(iv) Decarboxylation

Ans: (i) Acetylation refers to the process of introducing an acetyl group into a compound namely, the substitution of an acetyl group for an active hydrogen atom. Acetylation is usually carried out in presence of a base such as pyridine, dimethylaniline, etc.

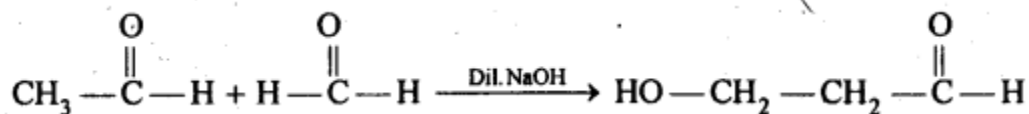
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(ii) **Cannizzaro reaction** : Aldehydes which do not contain an α -hydrogen atom, when treated with concentrated alkali solution undergo disproportionation, i.e., self oxidation reduction. As a result, one molecule of the aldehyde is reduced to the corresponding alcohol at the cost of the other which is oxidised to the corresponding carboxylic acid. This reaction is called Cannizzaro reaction.

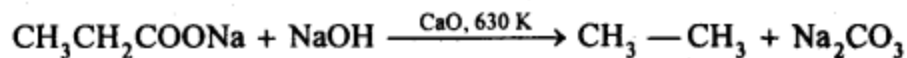


(iii) **Cross aldol condensation**: Aldol condensation between two different aldehydes is called cross aldol condensation. If both aldehydes contain α -hydrogens, It gives a mixture of four products.

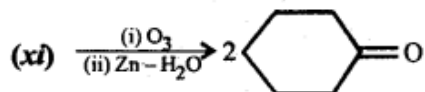
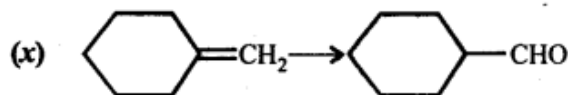
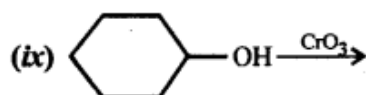
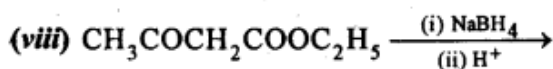
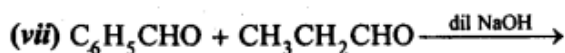
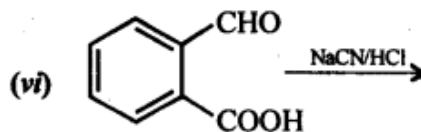
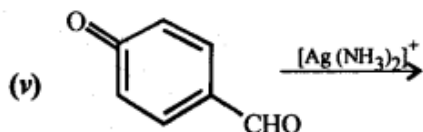
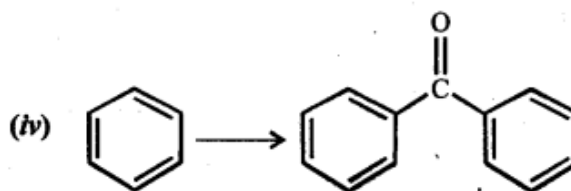
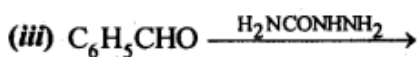
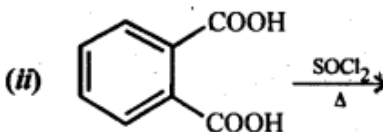
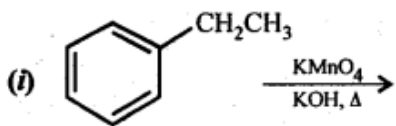


(iv) **Decarboxylation**: The process of removal of a molecule of CO_2 from a carboxylic acid is called decarboxylation. Sodium salts of carboxylic acids when heated with soda-lime undergoes decarboxylation to yield alkanes.

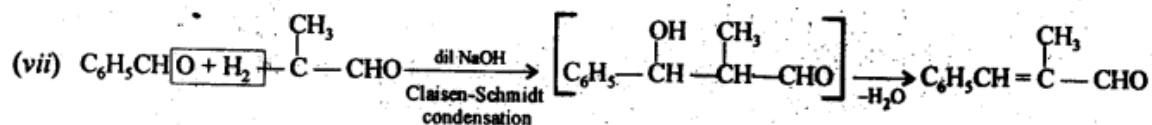
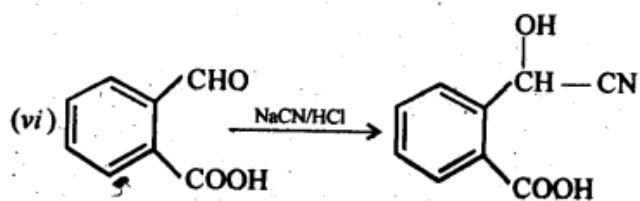
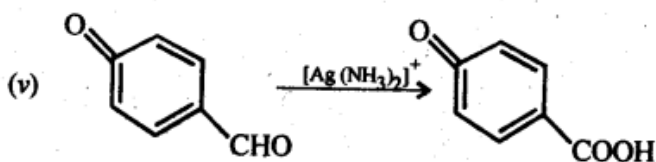
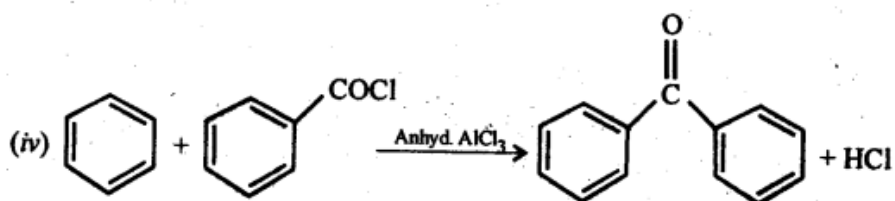
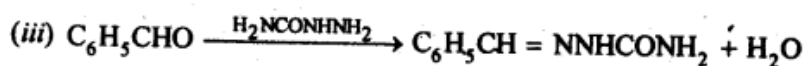
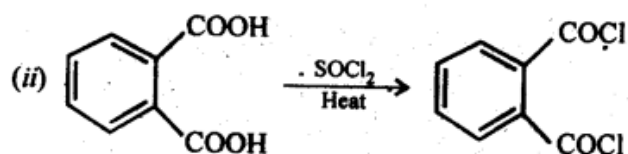
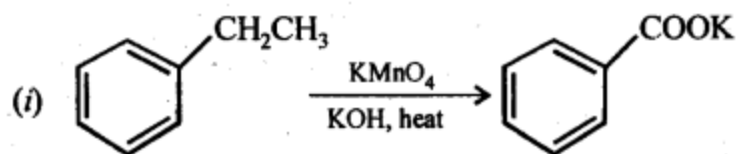
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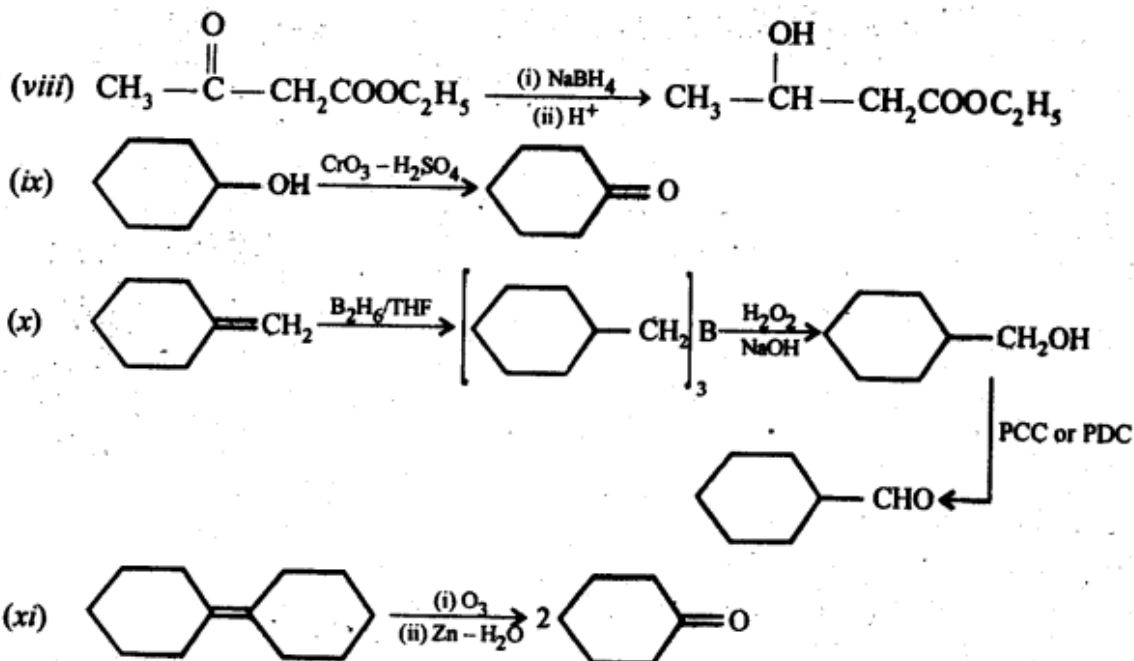


12.17. Complete each synthesis by giving missing starting material, reagent or products.



Ans:





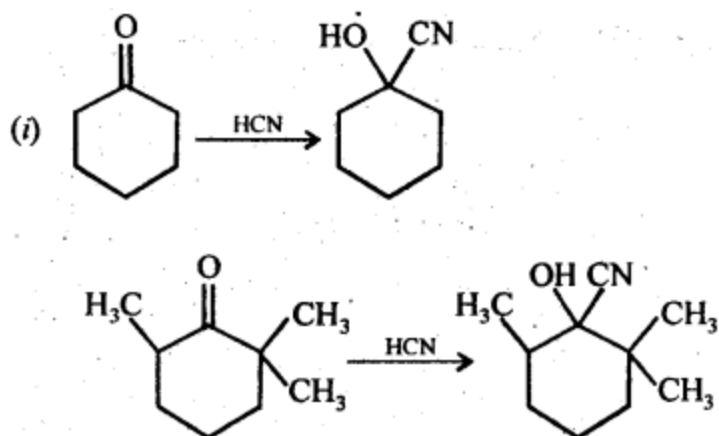
12.18. Give plausible explanation for each of the following:

(i) Cyclohexanone forms cyanohydrin in good yield but 2,2,6,6-tetramethylcyclohexanone does not

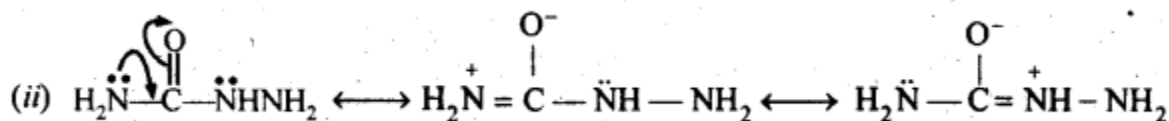
(ii) There are two $-\text{NH}_2$ groups in semicarbazide. However, only one is involved in the formation of semicarbazones.

(iii) During the preparation of esters from a carboxylic acid and an alcohol in the presence of an acid catalyst, the water or the ester should be removed as soon as it is formed.

Ans:



The yield of second reaction is very low because of the presence of three methyl groups at ex-positions with respect to the C = O, the nucleophilic attack by the CN^- ion does not occur due to steric hindrance. Since there is no such steric hindrance in cyclohexanone, therefore, nucleophilic attack by the CN^- ion occurs readily and hence cyclohexanone cyanohydrin is obtained in good yield.

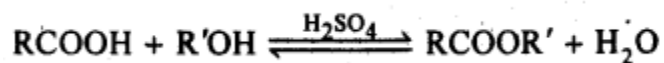


Although semicarbazide has two $-\text{NH}_2$ groups but one of them (i.e., which is directly attached to C = O) is involved in resonance as shown above. As a result, electron density on N of this $-\text{NH}_2$ group decreases and hence it does not act as a nucleophile. In contrast, the other $-\text{NH}_2$ group (i.e., attached to NH) is not involved in resonance and hence lone pair of electrons present

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on N atom of this -NH_2 group is available for nucleophilic attack on the $\text{C} = \text{O}$ group of aldehydes and ketones.'

(iii) The formation of esters from a carboxylic acid and an alcohol in presence of an acid catalyst is a reversible reaction.



Thus to shift the equilibrium in the forward direction, the water or the ester formed should be removed as fast as it is formed.

12.19. An organic compound contains 69.77% carbon, 11.63 % hydrogen and rest oxygen. The molecular mass of the compound is 86. It does not reduce Tottens' reagent but forms an addition compound with sodium hydrogensulphite and give positive iodoform test. On vigorous oxidation, it gives ethanoic and propanoic acid. Write the possible structure of the compound.

Ans:

C = 69.77%, H = 11.63%

$\therefore \text{O} = 100 - (69.77 + 11.63)\% = 18.6\%$

$\therefore \text{C} : \text{H} : \text{O} = \frac{69.77}{12} : \frac{11.63}{1} : \frac{18.6}{16} = 5.88 : 11.63 : 1.16 :: 5 : 10 : 1$

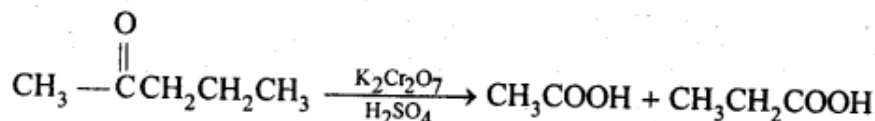
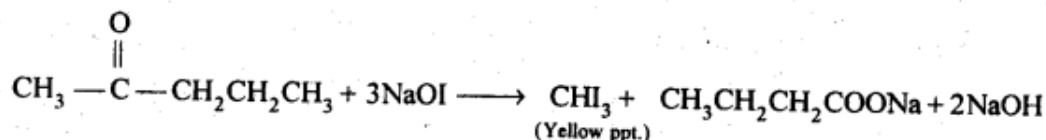
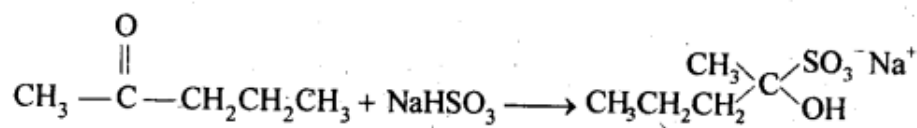
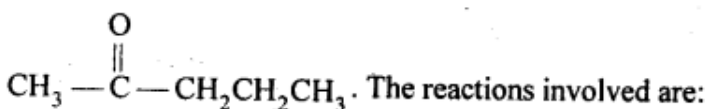
The empirical formula of the given compound = $\text{C}_5\text{H}_{10}\text{O}$

Empirical formula mass = $5 \times 12 + 10 \times 1 + 1 \times 16 = 86$

Molar mass = 86 (given)

$\therefore$ Molecular formula = $\text{C}_5\text{H}_{10}\text{O}$

Since the compound forms sodium hydrogen sulphite addition product, therefore, it must be either an aldehyde or methyl/cyclic ketone. Since the compound does not reduce Tollens' reagent, therefore, it cannot be an aldehyde. Since the compound gives positive iodoform test, therefore, the given compound is a methyl ketone. Since the given compound on vigorous oxidation gives a mixture of ethanoic acid and propanoic acid, therefore, the methyl ketone is pentan-2-one, i.e.,

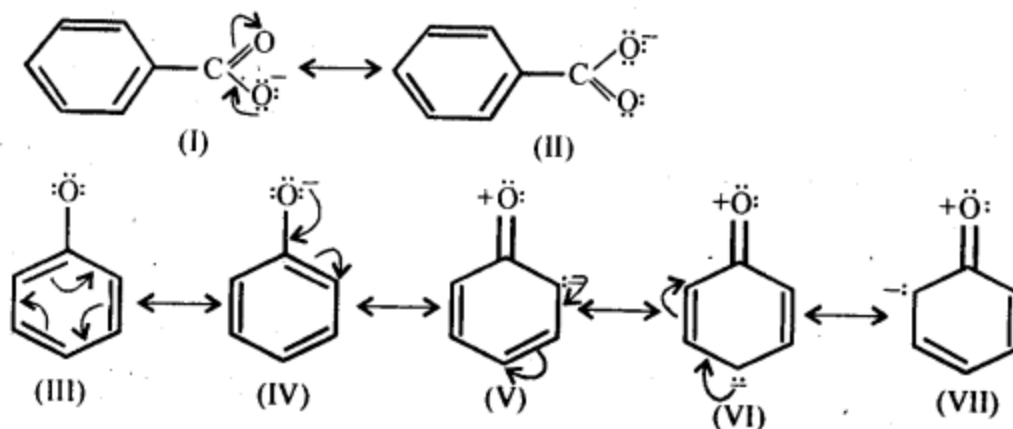


NCERT 12th Chemistry Chapter 12

12.20. Although phenoxide ion has more number of resonating structures than carboxylate ion, carboxylic acid is a stronger acid than phenol. Why?

Ans: Consider the resonating structures of carboxylate ion and phenoxide ion.

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In case of phenoxide ion, structures (V – VII) carry a negative charge on the less electronegative carbon atom. Therefore, their contribution towards the resonance stabilization of phenoxide ion is very small.

In structures I and II, (carboxylate ion), the negative charge is delocalized over two oxygen atoms while in structures III and IV, the negative charge on the oxygen atom remains localized only the electrons of the benzene ring are delocalized. Since delocalization of benzene electrons contributes little towards the stability of phenoxide ion therefore, carboxylate ion is much more resonance stabilized than phenoxide ion. Thus, the release of a proton from carboxylic acids is much easier than from phenols. In other words, carboxylic acids are stronger acids than phenols.

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- Chapter 5 : Surface
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Isolation of Elements
- Chapter 7 : The p Block
Elements
- Chapter 8 : The d and f
Block Elements
- Chapter 9 : Coordination
Compounds
- Chapter 10 : Haloalkanes
and Haloarenes
- Chapter 11 : Alcohols
Phenols and Ether
- Chapter 12 : Aldehydes
Ketones and Carboxylic
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